

Quiz #3y

Chemistry, 7th ed., Zumdahl & Zumdahl, section 2.8

Unless otherwise specified, each "blank" is worth 1 point.

Note that there is a Periodic Table of the Elements and some other possibly helpful information on the last page of this quiz.

PART 1

Identify each of the following compounds as ionic (**I**), molecular (**M**), or acid (**A**).

<u>Formula or Name</u>	<u>Compound Type</u> (I , M or A)
1. H_3PO_3	<u>A</u>
2. Cu_2SO_4	<u>I</u>
3. O_2	<u>M</u>
4. sulfur dioxide.....	<u>M</u>
5. Na_3PO_4	<u>I</u>

PART 2

Complete the following table. For **ionic compounds** or **acids ONLY**, identify the cation and anion.

<u>Formula</u>	<u>Cation</u>	<u>Anion</u>	<u>Name</u>
6. <u>XeF_6</u>	_____	_____	xenon hexafluoride
7. <u>CaBr_2</u>	<u>Ca^{2+}</u>	<u>Br^-</u>	<u>CALCIUM BROMIDE</u>
8. <u>K_2O</u>	<u>K^+</u>	<u>O^{2-}</u>	<u>POTASSIUM OXIDE</u>
9. <u>Mg_3N_2</u>	<u>Mg^{2+}</u>	<u>N^{3-}</u>	magnesium nitride
10. <u>SF_4</u>	_____	_____	<u>SULFUR TETRAFLUORIDE</u>
11. <u>HBr</u>	<u>H^+</u>	<u>Br^{1-}</u>	<u>HYDROBROMIC ACID</u> <u>HYDROGEN BROMIDE</u>
12. <u>Fe_2O_3</u>	<u>Fe^{3+}</u>	<u>O^{2-}</u>	<u>IRON (III) OXIDE</u>

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PART 1

Identify each of the following compounds as ionic (**I**), molecular (**M**), or acid (**A**).

<u>Formula or Name</u>	<u>Compound Type</u> (I , M or A)
1. CO ₂	<u>M</u>
2. K ₂ O	<u>I</u>
3. HC ₂ H ₃ O ₂	<u>A</u>
4. phosphorus pentachloride.....	<u>M</u>
5. Fe ₂ (CO ₃) ₃	<u>I</u>

PART 2

Complete the following table. For **ionic compounds** or **acids ONLY**, identify the cation and anion.

<u>Formula</u>	<u>Cation</u>	<u>Anion</u>	<u>Name</u>
6. <u>N₂O₄</u>	_____	_____	dinitrogen tetroxide
7. <u>MgI₂</u>	<u>Mg²⁺</u>	<u>I⁻</u>	<u>MgI₂</u>
8. <u>Na₃P</u>	Na ⁺	P ³⁻	<u>SODIUM PHOSPHIDE</u>
9. <u>Al₂O₃</u>	<u>Al³⁺</u>	<u>O²⁻</u>	aluminum oxide
10. <u>OF₂</u>	_____	_____	<u>OXYGEN DIFLUORIDE</u>
11. <u>HCl</u>	H ⁺	Cl ⁻	<u>HYDROCHLORIC / HYDROGEN ACID / CHLORIDE</u>
12. <u>PbI₄</u>	Pb ⁴⁺	I ⁻	<u>LEAD (IV) IODIDE</u>