

- Work with a classmate or two to complete these tables. *Check all your answers with at least one classmate before you turn this in.*
- HINTS: See sections 2.6-2.8 (especially tables 2.3-2.8) in your text. (*Chemistry 7th ed.*, Zumdahl & Zumdahl, Houghton Mifflin)
- Your instructor will help you if you can't agree on an answer or your group is stumped.

PART 1Identify each of the following compounds as ionic (**I**), molecular (**M**), or acid (**A**).

Formula or Name	Compound Type (I , M or A)
1. CO ₂	<u>M</u>
2. K ₂ O	<u>I</u>
3. tin(IV) bromide	<u>I</u>
4. HC ₂ H ₃ O ₂	<u>A</u>
5. H ₃ PO ₄	<u>A</u>
6. sodium chromate	<u>I</u>
7. N ₂ O	<u>M</u>
8. ammonium bromide	<u>I</u>
9. Cu ₂ SO ₄	<u>I</u>
10. hydrofluoric acid	<u>A</u>
11. MgH ₂	<u>I</u>
12. phosphorus pentachloride	<u>M</u>
13. CaI ₂	<u>I</u>
14. iron(III) carbonate	<u>I</u>
15. HClO	<u>A</u>
16. nickel(II) acetate	<u>I</u>
17. CCl ₄	<u>M</u>
18. barium carbonate	<u>I</u>
19. Fe ₂ (CO ₃) ₃	<u>I</u>
20. nitrous acid	<u>A</u>
21. Ni(C ₂ H ₃ O ₂) ₂	<u>I</u>
22. sulfur dioxide	<u>M</u>
23. Na ₃ PO ₄	<u>I</u>
24. oxygen difluoride	<u>M</u>

